
CHEMISTRY

9701/34

Paper 34 Advanced Practical Skills 2

May/June 2016

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Page 2	Mark Scheme	Syllabus	Paper
	Cambridge International AS/A Level – May/June 2016	9701	34

Question	Indicative material	Mark	Total
1 (a)	I Initial and final readings and titre value given for rough titre and initial and final readings for two (or more) accurate titrations (minimum of 2 x 2 box)	1	
	II Titre values recorded for accurate titrations and Appropriate headings for the accurate titration table and cm ³ units. <ul style="list-style-type: none"> initial / start burette reading / volume / value final / end burette and reading / volume / value titre or volume / FB3 and used / added unit: / cm³ or (cm³) or in cm³ or cm³ (for each heading) 	1	
	III All accurate burette readings are to the nearest 0.05 cm ³ . <i>Do not award this mark if:</i> <ul style="list-style-type: none"> 50(.00) is used as an initial burette reading more than one final burette reading is 50.(00) any burette reading is greater than 50.(00) 	1	
	IV There are two (or more) uncorrected, accurate titres within 0.10 cm ³ <ul style="list-style-type: none"> <i>Do not award this mark if, having performed two titres within 0.1 cm³, a further titration is performed which is more than 0.10 cm³ from the closer of the two initial titres, unless a further titration, within 0.10 cm³ of any other, has also been carried out.</i> <i>Do not award the mark if any “accurate” burette readings (apart from initial 0 cm³) are given to zero dp.</i> 	1	
	V, VI and VII Examiner rounds any accurate burette to the nearest 0.05 cm ³ , checks subtractions and then select the ‘best’ titres using the hierarchy: <ul style="list-style-type: none"> two (or more) accurate identical titres, <i>then</i> two (or more) accurate titres within 0.05 cm³, <i>then</i> two (or more) accurate titres within 0.10 cm³, <i>etc.</i> These best titres should be used to calculate the mean titre, expressed to nearest 0.01 cm ³ . Accuracy marks are awarded as shown. Award V, VI and VII for $\delta \leq 0.30$ (cm ³) Award V and VI for $0.30 \text{ cm}^3 < \delta \leq 0.60$ (cm ³) Award V for $0.60 \text{ cm}^3 < \delta \leq 1.00$ (cm ³)	3	[7]

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Question	Indicative material	Mark	Total
(b)	<p>Candidate must take the average of two (or more) titres that are within a total spread of not more than 0.20 cm^3. Working must be shown or ticks must be put next to the two (or more) accurate readings selected. The mean should be quoted to 2 dp, rounded to the nearest 0.01.</p> <p>Two special cases where the mean may not be to 2 dp:</p> <ul style="list-style-type: none"> Allow mean expressed to 3 dp only for 0.025 or 0.075 (e.g. 26.325) Allow mean if expressed to 1 dp and the mean is exactly correct. (e.g. 26.0 and 26.2 = 26.1 is allowed) (e.g. 26.0 and 26.1 = 26.1 is incorrect – should be 26.05.) Note: the candidate's mean will sometimes be marked as correct even if it is different from the mean calculated by the examiner for the purpose of assessing accuracy. 	1	[1]
(c) (i)	Correctly calculates $n(\text{HCl})$ used = $\frac{0.008 \times (\text{b})}{1000}$	1	[5]
(ii)	Correct use of factors $n(\text{borax}) = 0.5 \times \text{answer (i)}$ and	1	
(iii)	$n(\text{borax}) = 40 \times \text{answer (ii)}$		
(iv)	Correct use of (iii) and 15.5 $M_r = \frac{15.5}{(\text{iii})}$	1	
(v)	Correct expression $x = \frac{(\text{iv}) - 338}{(10.8)}$	1	
	All answers to 3 or 4 significant figures and if answer (v) is attempted it must be an integer	1	[5]
Question 2	[13]		

Page 4	Mark Scheme	Syllabus	Paper
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Question	Indicative material	Mark	Total
2 (a)	I Appropriate headings and units for the three balance readings <ul style="list-style-type: none"> (Mass of) crucible (and lid) (Mass of) crucible, (lid) and FB 4 (or “contents before heating”) (Mass of) crucible, (lid) and residue /MgO/ contents after heating / FB 4 after heating Unit covering all balance readings and subtracted values if shown: /g, (g), in g or g (for each heading) 	1	[5]
	II Masses recorded <ul style="list-style-type: none"> Mass of FB 4 used was claimed to be between 1.1–1.3 g All balance readings recorded to same number of decimal places (<i>at least one dp</i>) 	1	
	III Mass of FB 4 and of residue <ul style="list-style-type: none"> Mass of FB 4 used, correctly subtracted Mass of residue, correctly subtracted 	1	
	IV and V <ul style="list-style-type: none"> Use corrected values Examiner used corrected values and works out the ratio $\frac{\text{mass of FB 4}}{\text{mass of MgO}}$ to 1 dp for the candidate Accuracy marks are awarded as shown. Award IV if ratio between 1.4–2.5 Award V if ratio between 1.7–2.3	2	
(b) (i)	Correctly calculates $n(\text{MgO})$ <ul style="list-style-type: none"> $\frac{\text{mass of residue}}{40.3}$ <i>Answer must be expressed to 2, 3 or 4 significant figures</i> 	1	[5]
(ii)	Correct use of (i) and mass of FB 4 <ul style="list-style-type: none"> $n(\text{FB 4}) = \frac{\text{answer (i)}}{2}$ $M_r = \frac{\text{mass of FB 4 used}}{\text{no of moles of FB 4}}$ <i>An answer for M_r must be quoted to 2 or more significant figures</i> 	1	
(iii)	M_r calculated from A_r values in Periodic Table = 178.6	1	
(iv)	Correct expression shown $\frac{2.5}{100} \times M_r$ in (iii) (= 4.5, 4.47, 4.465) or expresses % difference of the two M_r values $= \frac{ (\text{iii}) - (\text{ii}) }{(\text{iii})} \times 100$ or (iii) $\times \frac{97.5}{100} / (\text{iii}) \times \frac{102.5}{100}$ to give range (= 174(.1) – 183(.1))	1	
	Makes a correct statement (support/does not support/yes/no) about the accuracy of the possible formula, explained by whether the experimental M_r value is close to the answer in (iii). <i>Numbers must be quoted or reference made to (ii) and (iii)</i>	1	[5]

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(c) (i)	Improvement Heat (crucible and residue) to constant mass <i>Accept a description of the procedure for the mark</i> or heat more strongly / to a higher temperature or heat for longer so more is decomposed If a 1 dp balance is used allow use a balance weighing to more dp and to reduce % error (in weighing) / give more precise mass	1	[4]
(ii)	Conclusion To find out whether the two experiments are reliable / consistent / concordant or If the experiments do not agree then carry out a 3rd / another expt or If one experiment was inaccurate because it gave a poor M_r value it can be ignored	1	
(iii)	Error is 0.005 g or 0.01 g (<i>if 2 dp balance was used</i>) (<i>If a 3 dp balance was used, error is 0.0005 or 0.001 g</i>) (<i>If a 1 dp balance was used, error is 0.05 or 0.1 g</i>)	1	
	% = $100 \times 2 \times \text{error} / \text{mass of FB 4}$ <i>Accept correct expression or correct answer to 2, 3 or 4 significant figures</i>	1	
Question 2			

[14]

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	FB 5 is BaCl ₂ ; FB 6 is AgNO ₃ ; FB 7 is H ₂ SO ₄ ; FB 8 is NaHCO ₃					
3 (a)	Three correct observations with Mg Three correct observations with NaOH Three correct observations with KI				3	
	<i>test</i>	FB 5	FB 6	FB 7		
	Mg	no reaction / no change (ignore few bubbles (on Mg) / gas produced slowly)	grey / black and ppt / solid / coating / deposit or Mg goes black / grey	fizzing / bubbling / effervescence		
	NaOH	no change / no reaction / no ppt / (solution) stays colourless <i>Allow white ppt / cloudy white</i> soluble in excess is CON	(dark) brown ppt soluble in excess is CON <i>Allow brown ppt as final colour after a paler shade of brown but not a different colour (e.g. green / yellow / red)</i>	no change / no reaction / (solution) remains colourless / no ppt or temperature rises <i>Allow heat produced but not exothermic</i>		
	KI	no change / no reaction / (solution) remains colourless <i>Accept pale yellow solution</i>	(pale) yellow ppt <i>Allow greenish yellow (but not yellowish green)</i> <i>Allow cream-yellow</i>	no change / no reaction / (solution) remains colourless <i>Accept pale yellow solution</i>		
	FB 5		white ppt	white ppt		
	(Mg)			gas / H ₂ (evolved) pops with lighted splint / spill		
Two correct observations with FB 5					1	
Correct hydrogen gas test					1	[5]

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(b) (i)	Conclusion and reason Cation is silver / Ag^+ (allow lead(II) / Pb^{2+}) and Gives yellow precipitate (with iodide ions) or AgI (PbI_2) produced	1	[4]
	(ii) Both conclusions about FB 5 are correct <ul style="list-style-type: none"> cation – barium / Ba^{2+} anion – chloride / Cl^- 	1	
	(iii) FB 7 is sulfuric acid / H_2SO_4	1	
	(iv) $\text{Mg} + 2\text{H}^+ \rightarrow \text{Mg}^{2+} + \text{H}_2$	1	
(c) (i)	Any two observations from: <ul style="list-style-type: none"> Condensation / water formed (inside test tube) Steam liberated / hissing / (water) vapour produced / steamy / misty fumes White residue / solid remains white / white solid formed 	1	[4]
	<u>Gas / CO_2</u> turns lime water milky / white precipitate / cloudy white / chalky	1	
	(ii) Any two observations from: <ul style="list-style-type: none"> Fizzing / bubbling / effervescence Reaction is rapid / vigorous Solid dissolves / colourless solution formed (at end) Temperature drops / it gets colder 	1	
	(iii) Both conclusions required <ul style="list-style-type: none"> cation – not known / not transition metal anion – carbonate / CO_3^{2-} or hydrogencarbonate / HCO_3^- 	1	
Question 3			[13]