
CHEMISTRY

9701/34

Paper 3 Advanced Practical Skills 2

October/November 2017

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Question	Answer	Marks
1(a)	I 5 (or more) experiments completed and Table to show Volume of FB 1 , Volume of water, Time and Rate	1
	II Correct units for all data Volume: in cm^3 or cm^3 or (cm^3) or cm^3 by each volume Time: /s or (s) or s by each time (<i>not sec or seconds but allow 'in seconds'</i>) Rate: /s ⁻¹ or (s ⁻¹) or s ⁻¹ by each rate	1
	III All times recorded to nearest second (minimum of 3 times)	1
	IV Two additional experiments with volume FB 1 not less than 10 cm^3 , not more than 40 cm^3 and no volume $\leq 2 \text{ cm}^3$ close to another volume.	1
	V Volumes of water chosen so that FB1 + water = 40 cm^3 for additional experiments carried out.	1
	VI Correctly calculates rate for all experiments and shown to 2 – 4 sf.	1
	VII Award if all candidate's times increase with decrease in volume of FB 1 .	1
	VIII Award if candidate's time to nearest second for Experiment 2 is within 10% of the supervisor's result	1
	IX Award if candidate's (time for FB 1 = 20)/(time FB 1 = 40) is between 1.90 and 2.40	1
	X Award if candidate's (time for FB 1 = 20)/(time FB 1 = 40) is between 2.00 and 2.30	1
1(b)	Linear scales that cover more than half the space in both directions and axes labelled correctly (allow the correct unit as the label)	1
	Points plotted correctly. Points must be within half a small square of the correct position, if the point should be on a line it must be on the line and if it should not be on the line it must not be so.	1
	Line of best fit drawn which ignores anomalous results identified by the candidate	1

Question	Answer	Marks
1(c)	Correct line drawn within 1 small square (horizontal line must be shown and some mark shown at 8).	1
	Correctly calculates = $1000/\text{rate}$ (to 2 – 4 sf or a whole number of seconds).	1
1(d)(i)	The print (on the insert) would take longer to disappear	1
	The liquid would be less deep	1
1(d)(ii)	The reaction time would be longer / reaction is slower / rate is less	1
	Accuracy improved because the percentage error in time less OR Accuracy not improved because more difficult to judge when print disappeared	1
1(e)	Expression % = $(1 / \text{Reaction time Experiment 1}) \times 100$ OR $(0.5 / \text{Reaction time Experiment 1}) \times 100$	1
1(f)	Keep volume thiosulfate / FB1 constant and vary volume acid / FB 2	1
	Keep total volume FB 2 + water constant	1
	Keep temperature constant / use same (shape) reaction vessel / use same printed sheet / carry out 5 (or more) expts with different volumes HCl / FB 2	1
1(g)(i)	Straight line through origin (with positive gradient)	1
1(g)(ii)	Straight horizontal line	1

Question	Answer				Marks
FB 3 is NaOH(aq), FB 4 is NH ₃ (aq), FB 5 is MgCl ₂ (aq), FB 6 is CuCl ₂ (aq), FB 7 is Na ₂ S ₂ O ₃ (aq), FB 8 is Na ₂ S ₂ O ₈ (aq), FB 9 is Na ₂ SO ₄ (aq).					
2(a)(i)		FB 4	FB 5	FB 6	3
	FB 3	No reaction / no change / solution remains colourless	White ppt	(Pale / light) blue ppt	
	FB 4		White ppt	Dark / deep blue solution / (pale/light) blue ppt	
	FB 5			No reaction / no change	
	6 correct boxes = 3 marks, 4 or 5 correct boxes = 2 marks, 2 or 3 correct boxes = 1 mark.				
2(a)(ii)	OH ⁻ / hydroxide				1
2(a)(iii)	Named indicator eg red litmus ('red' could be in the results) or formula / named (aqueous) salt that gives insoluble hydroxides				1
	Positive result for alkali				1
2(a)(iv)	Two of Mg ²⁺ , Zn ²⁺ , Al ³⁺ , Ca ²⁺ , Ba ²⁺				1
2(a)(v)	Test to distinguish ions in (iv)				1
	Result of test and appropriate conclusion				1

Question	Answer				Marks
2(b)(i)		FB 7	FB 8	FB 9	3
	KI	No reaction / no change / solution remains colourless	Yellow / brown colour	No reaction / no change / solution remains colourless	
	starch		then blue-black / black / dark blue		
	I ₂	Decolourises	No reaction	No reaction / (stays) yellow / brown	
	Ba ²⁺	No reaction no change / solution remains colourless / no ppt	(ignore responses here)	White ppt	
	9 correct boxes = 3 marks 6, 7 or 8 correct boxes = 2 marks 3, 4 or 5 correct boxes = 1 mark				

Question	Answer	Marks
2(b)(ii)	SO_4^{2-} or SO_3^{2-} (both needed)	1
2(b)(iii)	Add suitable named acid to FB 9 and $\text{Ba}(\text{NO}_3)_2 / \text{BaCl}_2$ ppt or Add (acidified aqueous) potassium manganate(VII) / KMnO_4 to FB 9 or Add named acid and test (any) gas evolved with (acidified aqueous) potassium manganate(VII)	1
	Anion present: SO_4^{2-} and No effect of acid on (white) ppt or (Solution) turns purple / purple not decolourised or No bubbles / manganate(VII) paper remains purple / blue litmus remains blue	1