

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**

International General Certificate of Secondary Education

**MARK SCHEME for the November 2003 question papers****0620 CHEMISTRY**

<b>0620/01</b>	<b>Paper 1 (Multiple Choice), maximum mark 40</b>
<b>0620/02</b>	<b>Paper 2 (Core), maximum mark 80</b>
<b>0620/03</b>	<b>Paper 3 (Extended), maximum mark 80</b>
<b>0620/05</b>	<b>Paper 5 (Practical), maximum mark 40</b>
<b>0620/06</b>	<b>Paper 6 (Alternative to Practical), maximum mark 60</b>

These mark schemes are published as an aid to teachers and students, to indicate the requirements of the examination. They show the basis on which Examiners were initially instructed to award marks. They do not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the November 2003 question papers for most IGCSE and GCE Advanced Level syllabuses.



**Grade thresholds** taken for Syllabus 0620 (Chemistry) in the November 2003 examination.

	maximum mark available	minimum mark required for grade:			
		A	C	E	F
Component 1	40	-	28	22	18
Component 2	80	-	58	45	36
Component 3	80	46	28	-	-
Component 5	40	28	22	17	13
Component 6	60	44	34	25	19

The threshold (minimum mark) for B is set halfway between those for Grades A and C.  
The threshold (minimum mark) for D is set halfway between those for Grades C and E.  
The threshold (minimum mark) for G is set as many marks below the F threshold as the E threshold is above it.

Grade A\* does not exist at the level of an individual component.

November 2003

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 40**

**SYLLABUS/COMPONENT: 0620/01**

**CHEMISTRY**

**(Multiple Choice)**



<b>Page 1</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>IGCSE – NOVEMBER 2003</b>	<b>0620</b>	<b>1</b>

<i>Question Number</i>	<i>Key</i>	<i>Question Number</i>	<i>Key</i>
1	A	21	D
2	D	22	A
3	B	23	C
4	C	24	A
5	B	25	B
6	D	26	B
7	A	27	B
8	C	28	A
9	B	29	D
10	B	30	D
11	B	31	D
12	D	32	B
13	A	33	A
14	D	34	C
15	D	35	B
16	B	36	A
17	C	37	C
18	C	38	A
19	C	39	B
20	C	40	B

November 2003

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 80**

**SYLLABUS/COMPONENT: 0620/02**

**CHEMISTRY**  
Core



Page 1	Mark Scheme	Syllabus	Paper
	IGCSE – NOVEMBER 2003	0620	2

- 1 (a) (i) B [1]
- (ii) A [1]
- (iii) C [1]
- (iv) D [1]
- (b) (i) distillation [1]  
ALLOW: fractional distillation
- (ii) chromatography [1]
- (c) (i) fuel gas [1]
- (ii) paraffin: any one of:  
oil stoves/heaters/for heating; aircraft fuel; [1]  
ALLOW: for lamps/for lighting/for cooking  
bitumen: any one of:  
road surfaces; ALLOW: for roads  
roofing tar; NOT: 'tar' without qualification  
in/for electrical cables; NOT: electrical cables [1]
- (d) hydrocarbons [1]
- (e) correct structure with correct pairings of dots and crosses [2]  
(correct structure with only dots or only crosses/random dots/crosses = 1)  
IGNORE: lack of inner electron shell of carbon
- (f) alkane(s) [1]
- Total = 13**
- 2 (a) respiration [1]
- (b) (i) lighted splint/put mouth of test tube of hydrogen in flame; [2]  
pops/explosion
- (ii) makes explosive mixture [1]
- (iii) (red) litmus paper/universal indicator paper/pH paper; [2]  
turns blue  
ALLOW: HCl gas/HCl on glass rod; white fumes
- (c) correct displayed/geometric formula for ethanoic acid [1]  
(all bonds must be shown)
- (d) (i) 2H<sub>2</sub> [1]
- (ii) fuel/making ethyne/making carbon black/making synthesis gas/ [1]  
making methanol  
NOT: natural gas  
NOT: cooking

Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – NOVEMBER 2003	0620	2

(iii) 1<sup>st</sup> and 3<sup>rd</sup> boxes ticked [2]  
1 box correct = 1 mark

(e) (i) zinc [1]

(ii) iron/nickel [1]  
ALLOW: zinc

(iii) lead [1]

(iv) calcium carbonate [1]

(v) aluminium [1]

**Total = 16**

3 (a) In iron making  
ALLOW: in blast furnace/for neutralising acid soils or acidic lakes etc./for  
building/making cement OR concrete/hard core/road foundations [1]  
NOT: removing impurities from iron ore  
NOT: purification of water

(b)  $C + O_2 \rightarrow CO_2$   
correct formula for oxygen;  
correct formula for carbon dioxide [2]  
(-1 per other error)  
ALLOW:  $2C + O_2 \rightarrow 2CO$  (2 marks)

(c) exothermic [1]  
NOT: combustion

(d) calcium oxide;  
carbon dioxide [2]  
NOT: symbols

(e) (i) 2 (HCl) [1]

(ii) limewater;  
turns milky/cloudy [2]

(f) (i) oxidation;  
the carbon has gained oxygen/oxidation number of carbon has  
increased/carbon has lost electrons [2]  
(the answer must refer to the carbon)  
NOT: carbon gets oxidised

(ii) blowtorches/welding/cutting metals [1]  
ALLOW: to make (monomers for) neoprene/synthetic rubber  
NOT: other organic syntheses

**Total = 12**

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – NOVEMBER 2003	0620	2

- 4 (a) halogen(s) [1]
- (b) (i) (atoms with same atomic number) but different mass number/different numbers of neutrons/different nucleon number [1]  
NOT: atoms with different atomic masses
- (ii) 35 + 35 [1]  
44; 46 [2]  
35 + 35 [1]
- (c) (i) chlorine more reactive (than bromine)/higher in the reactivity series (than bromine) (or reverse argument) [1]  
ALLOW: it is more reactive  
NOT: chlorine higher in the table
- (ii) potassium bromide + chlorine → potassium chloride + bromine [1]  
ALLOW: completely correct symbol equation
- (d) (i) 3.5 [1]  
ALLOW: 3.3- 3.5
- (ii) pH 3 [1]
- (iii) pH 7 [1]
- (e) bromine (water) decolourised/goes from red-brown/orange/brown to colourless [1]  
ALLOW: it is decolourised  
NOT: incorrect colours to colourless

**Total = 12**

- 5 (a) 5 (O<sub>2</sub>) [1]
- (b) anhydrous/white copper sulphate; [2]  
turns blue  
OR  
anhydrous/blue cobalt chloride;  
turns pink  
NOT: boiling point 100°C
- (c) 1<sup>st</sup> and 2<sup>nd</sup> boxes ticked [2]
- (d) (i) carbon monoxide [1]
- (ii) incomplete combustion of the fuel/gas/burning in limited amount of oxygen/air [1]  
NOT: incomplete burning  
NOT: lack of air
- (e) (i) gas [1]



Page 4	Mark Scheme	Syllabus	Paper
	IGCSE – NOVEMBER 2003	0620	2

- (ii) coal [1]
- (iii) any two of:  
erodes buildings containing (calcium) carbonate OR erodes/corrodes metals in buildings;  
NOT: corrodes (calcium) carbonate  
kills (small) animals in water OWTTE/kills pond life;  
(NOT: kills animals)  
damages trees/plants/causes leaf burn/damages plant roots; [2]  
ALLOW: kills plants  
NOT: causes breathing difficulties  
NOT: destroys buildings/wildlife/plants/animals
- (iv) white;  
precipitate/solid [2]
- Total = 13**
- 6 (a) aluminium high in reactivity series/too reactive [1]  
ALLOW: aluminium higher in reactivity series than carbon  
ALLOW: carbon will not reduce aluminium oxide
- (b) electrical heating [1]  
NOT: heating
- (c) conducts electricity/ [1]  
ALLOW: good conductor  
NOT: has high melting point/inert/unreactive
- (d) cathode [1]
- (e) saves energy/too much energy required to melt aluminium oxide;  
ALLOW: too much heat required/electricity OR heat is expensive  
NOT: unqualified 'expensive':  
will not melt the steel casing [2]  
ALLOW: melting point is higher than steel  
NOT: melting point too high
- (f) any two of :  
oxygen reacts with the carbon/graphite/(positive) electrode/anode gets oxidised;  
carbon dioxide formed;  
carbon electrodes/anodes decrease in size/get eroded away [2]  
ALLOW: anodes get eaten away/wear away  
NOT: anodes dissolve
- (g)  $3 e^-$  [1]  
ALLOW  $3e$
- (h) positive ions attracted to negative electrode/positive charges attracted to negative/aluminium has oppositely charged ions to the negative electrode;  
ALLOW: aluminium ions are positive [1]

<b>Page 5</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>IGCSE – NOVEMBER 2003</b>	<b>0620</b>	<b>2</b>

(i) 60% [1]

(ii) 3 from:  
malleable;  
ductile;  
sonorous;  
shiny;  
conduct heat;  
conduct electricity [3]  
ALLOW: flexible/bendy  
NOT: high melting/boiling points/high densities

**Total = 14**

November 2003

**INTERNATIONAL GCSE**

**MARK SCHEME**

**MAXIMUM MARK: 80**

**SYLLABUS/COMPONENT: 0620/03**

**CHEMISTRY**  
**Extended**



- An incorrectly written symbol, e.g. NA **or** CL, should be penalised once in a question.

In the mark scheme if a word **or** phrase is underlined it(**or** an equivalent) is required for the award of the mark.

(.....) is used to denote material that is not specifically required.

**OR** designates alternative and independent ways of gaining the marks for the question.

**or** indicates different ways of gaining the same mark.

**cond** indicates that the award of this mark is conditional upon a previous mark being gained.

- Unusual responses which include correct Chemistry that answers the question should always be rewarded - even if they are not mentioned in the marking scheme.
- All the candidate's work must show evidence of being marked by the examiner.

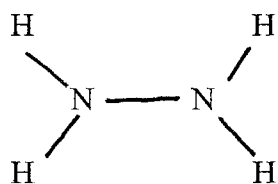
- 1 (a) (i) different boiling points [1]
- (ii) methane **or** water **or** petroleum **or** named petroleum fraction **or** alkane [2]
- Any TWO** [2]
- (b) (i) volume decrease for forward reaction **or** fewer moles of gas on products side [1]
- favoured by increase in pressure [1]
- or** increase in pressure moves position of equilibrium to right
- (ii) increase [1]
- exothermic reaction favoured by lower temperature [1]
- (iii) 300 to 600 °C
- 1:3 volume ratio
- iron (catalyst)
- 150 to 300 atm
- Any TWO** [2]
- (c) (i) proton [2]
- hydrogen ion **or** H<sup>+</sup> **ONLY** [1]
- (ii) correct equation molecular **or** ionic [1]
- $\text{NH}_3 + \text{HCl} = \text{NH}_4\text{Cl}$
- $\text{NH}_3 + \text{H}^+ = \text{NH}_4^+$  accept  $\text{NH}_4\text{OH}$
- (d) measure pH **or** add universal indicator **or** pH meter [1]
- ammonia has lower pH if numerical values given

must be appropriate that is above 7 with ammonia having the lower value **or** correct colours, green and blue are acceptable [1]

**OR** measure conductivity [1]

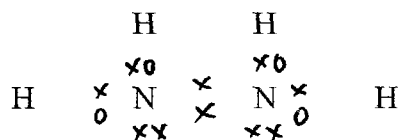
ammonia has poorer conductivity [1]

(e) (i) correct structural formula [1]



(ii) 8e around nitrogen [1]

2e around each hydrogen [1]



TOTAL = 17

2 (a) (i) 40 [1]

80 **or** 40 [1]

1 [1]

(ii) particles have more energy **or** moving faster [1]  
collide more frequently

**or** collide with more energy [1]

(iii) greater surface area [1]

(iv) flour mills **or** coal mines **or** metal powders [1]

**or** fireworks **or** gunpowder

(b) (i) collect and measure volume of oxygen [1]

**or** mass **or** count bubbles

time [1]

(ii) measure rate in different light levels and comment [1]

accept if dark no reaction

(c) (i) +6O<sub>2</sub> [2]

not balanced that is just O<sub>2</sub> **ONLY** [1]

(ii) linkage ---O---- [1]

chain [1]

minimum to be accepted

TOTAL = 14

- 3 (a) (i) heat **or** roast [1]  
in air [1]
- (ii) Either correct equation [2]  
 $\text{ZnO} + \text{C} = \text{Zn} + \text{CO}$   
 $2\text{ZnO} + \text{C} = 2\text{Zn} + \text{CO}_2$   
Not balanced **ONLY** [1]  
NOT carbon monoxide as a reductant
- (iii) bp of lead above  $1400\text{ }^{\circ}\text{C}$  it remains  
bp of zinc below  $1400\text{ }^{\circ}\text{C}$   
boils away **or** forms vapour  
**Any TWO** [2]
- OR** lead does not boil [1]  
zinc boils [1]
- (b) (i) making brass **or** any zinc containing alloy **or** galvanising [1]  
**or** sacrificial protection **or** batteries **or** roofs
- (ii) lattice **or** layers of (positive) ions  
delocalised **or** free **or** mobile electrons  
layers/atoms/particles can slip [3]
- (iii) different size atom **NOT** shape [1]  
prevents layers from moving [1]
- (c) (i) one involving lead – change 2 [1]  
**cond** because electrons are gained [1]  
**or** oxidation number less
- (ii) correct equation [2]  
 $\text{Zn} + 2\text{Ag}^+ = 2\text{Ag} + \text{Zn}^{2+}$   
not balanced **ONLY** [1]

TOTAL = 16

- 4 (a) (i) in which something dissolves [1]  
(ii) correct formula [1]  
 $\text{CH}_3\text{COOC}_2\text{H}_5$  **or** full structural formula

**NOT** C<sub>4</sub>H<sub>8</sub>O<sub>2</sub>

(iii) steam **or** water **or** hydration [1]  
heat **or** catalyst [1]

**OR** bubble into (concentrated) sulphuric acid [1]  
add water [1]

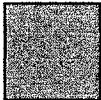
oxidised [1]  
by air **or** dichromate **or** manganate(VII) [1]

(iv) ethanoic acid and butanol [1]

(b) (i) CH<sub>2</sub>OH [1]  
CHOH  
CH<sub>2</sub>OH

(ii) soap **or** detergent [1]

(c) (i) polyester **or** condensation polymer **NOT** terylene [1]

(ii) HOOC -  -COOH [1]

HO -  -OH [1]

If wrong way around [1] Point of attachment of functional group to “box” not important

(d) (i) protein **or** poly peptide **or** polyamide [1]

(ii) peptide **or** amide [1]

(iii) amino acids are colourless **or** become visible/coloured **or** to develop it [1]

(iv) using colour **or** from position **ONLY** [1]

**OR** discussion of R<sub>f</sub> [2]

**OR** compare with known amino acids [2]

TOTAL = 17

5 (a) (i) preserve food **or** sterilising [1]

(ii) making paper [1]

- (b) (i) making sulphuric acid **or** Contact Process [1]  
(ii) oxygen [1]  
(iii) vanadium oxide as catalyst (ignore oxidation state)  
400 to 500 °C  
pressure less than 10 atm  
**Any TWO** [2]
- (c) (i) pink **or** purple [1]  
colourless **NOT** clear [1]  
(ii) barium sulphate [1]  
**cond** bromine oxidises **or** reacts with [1]  
sulphur dioxide to form sulphate ion [1]
- (d) the number of moles of SO<sub>2</sub> in the mixture = 0.125  
the number of moles of Cl<sub>2</sub> in the mixture = 0.2  
**cond** reagent was not in excess? SO<sub>2</sub>  
**cond** moles of SO<sub>2</sub>Cl<sub>2</sub> formed = 0.125  
**cond** the mass of sulphuryl chloride formed = 16.9g [5]

TOTAL = 16

TOTAL for PAPER = 80



**CAMBRIDGE**  
INTERNATIONAL EXAMINATIONS

November 2003

**INTERNATIONAL GCSE**

MARK SCHEME

MAXIMUM MARK: 40

SYLLABUS/COMPONENT: 0620/05

CHEMISTRY  
Practical



0620/05 MARK SCHEME DETAILS  
November 2003

Question Number			Part Mark	
1		Times recorded in seconds in table	(1)	
		Times increasing in magnitude	(1)	
		Comparable to Supervisor's results	(2)	4
	(a)	Suitable scale for time/s	(1)	
		Points plotted correctly	(2)	
		-1 for each incorrect		
		Smooth line graph	(1)	4
	(b)	colourless	(1)	
		to blue/purple	(1)	2
	(c)	Estimate read from graph	(1)	
		Unit	(1)	2
		Indication correct on graph	(1)	1
	(d) (i)	Experiment 1	(1)	1
		(ii) Greatest concentration of bromate	(1)	
	therefore more collisions	(1)	2	
(e) (i)	Two sources of error	(2)		
	e.g. inaccurate measurement of named liquid into beaker/stopping timer at same colour level/timing problem – <u>not</u> use of timer		2	
	(ii) Two improvements	(2)		
	e.g. use a burette/use a colorimeter/read and average		2	
			<b>Sub-total = 20</b>	
2	(a)	yellow/brown/orange	(1)	1
	(b) (i)	orange/brown	(1)	
		precipitate	(1)	
		remains/insoluble in excess	(1)	3
		(ii) Litmus → blue/indicator pH > 7	(1)	
		smell of gas	(1)	2
	(c)	orange/brown	(1)	
		precipitate	(1)	
	remains/insoluble in excess	(1)	3	

(d)	fizz/bubbles etc	(1)	
	lighted splint	(1)	
	pops	(1)	3
	green	(1)	
	precipitate	(1)	
			2
(e)	white	(1)	
	precipitate	(1)	
			2
(f)	iron(III) = 2 marks, iron only = 1 mark	(2)	
	ammonium	(1)	
	sulphate	(1)	
	or correct formulae		4

**Sub-total = 20**

**Total for paper = 40**

November 2003

INTERNATIONAL GCSE

MARK SCHEME

MAXIMUM MARK: 60

SYLLABUS/COMPONENT: 0620/06

CHEMISTRY  
Alternative to Practical



Page 1	Mark Scheme	Syllabus	Paper
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Question Number	Question (Including any Source Details)	Part Mark
1 (a)	Boxes labelled clockwise: Condenser (1) Beaker (1) <u>Fractionating</u> column (1)	3
(b)	↑ underneath flask (1)	1
(c)	Fractional (1) distillation (1)	2
2 (a)	Larger surface area (1) Quicker to extract colour/more colour extracted (1) <u>not</u> easier/faster	2
(b)	Reference to ethanol (1)	1
(c)	Reference to flammability of ethanol (1)	1
(d)	To prevent loss of solvent (1) <u>not</u> splash/evaporation	1
(e)	<u>Pour</u> off liquid (1)	1
(f)	Chromatography (1) Apply orange concentrate (1) to paper (1) Use of solvent (1) Description of elution (1) Result of experiment (1)  Max 5 – all marks could be obtained from a suitable diagram	5
3	Table. Times read correctly: 4s (1) 8s (1) 14s (1) 30s (1) 82s (1)	5
(a)	Points plotted correctly (3) (-1 for each incorrect) Smooth line graph (1)	4
(b)	Read from graph – should be $\approx$ 48 (1) <u><math>\pm</math></u> (1) Indication on graph (1)	2 1
(c) (i)	Experiment 1 (1)	1
(ii)	Greatest concentration/amount of bromate (1) Therefore more <u>collisions</u> (1)	2

Page 2	Mark Scheme	Syllabus	Paper
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Question Number	Question (Including any Source Details)	Part Mark
(d) (i)	Two errors: e.g. use of m cylinder inaccurate (1)/use of timer (1)/detecting when cross not visible	2
(ii)	Improvements: e.g. use of burette (1)/use of computer data logging (1)/use of colourimeter (1) insulate repeat and average	2
4 (b) (i)	Orange/brown (1) Precipitate (1) No change in excess (1)	3
(c)	Orange/brown precipitate (1) No change in excess (1)	2
(f) (i)	Hydrogen (1)	1
(ii)	Reduction/redox/displacement (1) iron (II) formed (1)	2
(g)	Cation – ammonium (1) Anion – sulphate (1)	2
5 (a)	Sodium hydroxide (1)	1
(b)	Ammonium sulphate (1)	1
(c)	Bunsen burner (1)	1
(d)	Reference to reaction (1)	1
(e)	Gas jar wrong way up (1) Gas is less dense than air (1)	2
	Tubes in flask should be evened (1) Liquid would be transferred to gas jar (1)	2
	Also credit in (c)	
6	Weigh coal sample (1) same amount <u>Burn</u> coal (1) Pass gas or diagram to show (1) Through acid/dichromate (1) Use of timer (1) Record time for colour change (1) Repeat/compare with other samples (1)	
	Max 6	6
		<b>Total 60</b>