

Centre Number

Candidate Number

Name

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education**CHEMISTRY****0620/02**

Paper 2 (Core)

October/November 2005

**1 hour 15 minutes**Candidates answer on the Question Paper.  
No Additional Materials required.**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.  
Write in dark blue or black pen in the spaces provided on the Question Paper.  
You may use a pencil for any diagrams, graphs or rough working.  
Do not use staples, paper clips, highlighters, glue or correction fluid.

Answer **all** questions.

The number of marks is given in brackets [ ] at the end of each question or part question.  
A copy of the Periodic Table is printed on page 16.

**For Examiner's Use**

1	
2	
3	
4	
5	
6	
<b>Total</b>	

This document consists of **16** printed pages.

1 The diagram shows part of the Periodic Table.

				He
C	N	O	F	Ne
		S	Cl	Ar
			Br	Kr

(a) Answer these questions using only the elements shown in the diagram.

Write down the symbol for an element which

- (i) has five electrons in its outer shell, ..... [1]
- (ii) has diatomic molecules, ..... [1]
- (iii) reacts with sodium to form sodium bromide, ..... [1]
- (iv) is a noble gas, ..... [1]
- (v) has a giant covalent structure, ..... [1]
- (vi) has a lower proton number than fluorine, ..... [1]
- (vii) is the most abundant gas in the air. .... [1]

(b) Write down a use for each of the following elements.

- (i) argon  
..... [1]
- (ii) helium  
..... [1]
- (iii) oxygen  
..... [1]

(c) (i) Draw a diagram to show the electronic structure of argon.

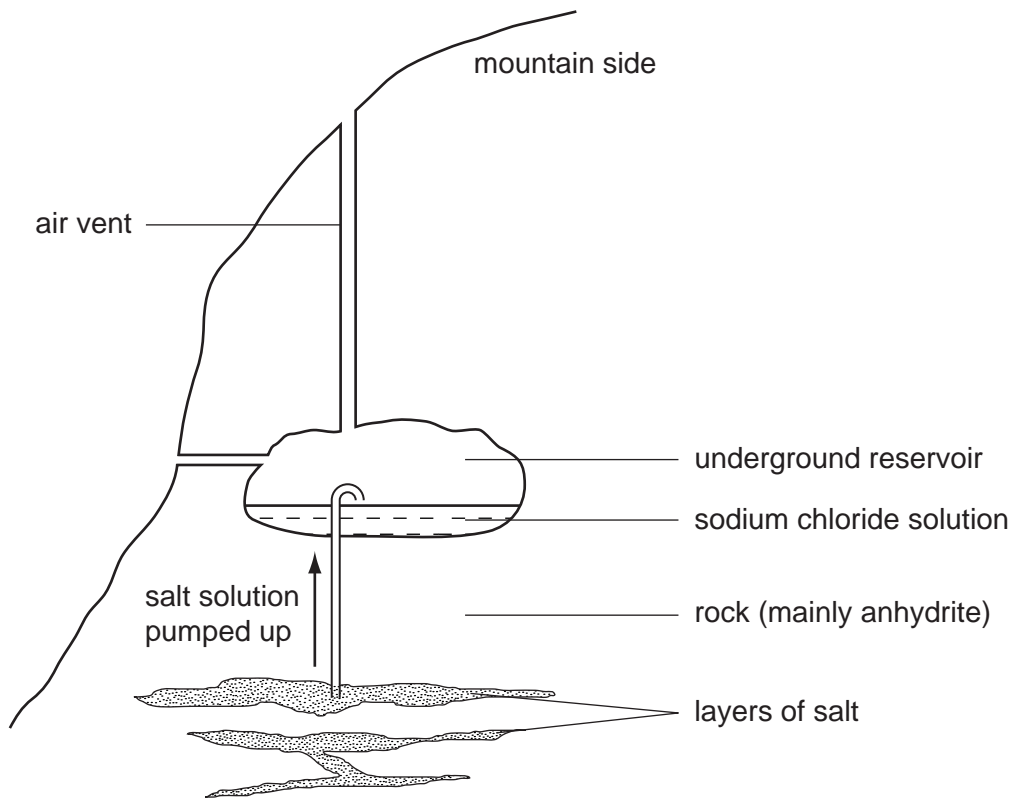
[2]

(ii) Why is argon very unreactive?

[1]

.....

2 The diagram shows the salt mines at Bex in Switzerland.



The salt is dissolved by water from underground springs and then pumped up to a reservoir where it is stored as a solution.

(a) Write the chemical formula for sodium chloride.

..... [1]

(b) Suggest how solid sodium chloride is obtained from the sodium chloride solution.

..... [1]

- (c) Sodium chloride has an ionic giant structure.  
Which one of the following best describes an aqueous solution of sodium chloride?  
Tick one box.

a mixture of sodium ions and chlorine molecules in water

a mixture of sodium and chlorine atoms in water

a mixture of sodium and chloride ions in water

a mixture of sodium, chloride, oxide and hydrogen ions

[1]

- (d) Describe a test for chloride ions.

test .....

result ..... [2]

- (e) The rock surrounding the layers of salt is anhydrite.  
Pure anhydrite has the chemical formula  $\text{CaSO}_4$ .

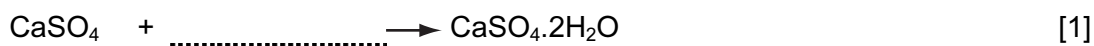
- (i) State the name of the chemical found in anhydrite.

..... [1]

- (ii) Calculate the relative formula mass of the chemical in pure anhydrite.

[1]

- (iii) When anhydrite reacts with water, gypsum ( $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ ) is formed.  
Complete the equation for this reaction.



- (iv) Which one of the following describes this reaction?  
Put a ring around the correct answer.

combustion    fermentation    hydration    oxidation    reduction    [1]

- (v) The chemical in anhydrite can be made by reacting calcium hydroxide with sulphuric acid.  
Complete the balanced equation for this reaction.



- (vi) The spring water running through the rocks changes anhydrite into gypsum.  
This reaction is exothermic.  
Use this information to explain why the temperature of the mine never falls below 17°C even in cold winters.

.....  
..... [1]

- (f) The air inside the mine contains 19% oxygen.  
Which one of the following best describes the oxygen level inside the mine compared with that outside the mine?  
Tick one box.

the level of oxygen inside the mine is higher

the level of oxygen is the same

the level of oxygen is about a quarter of that of the outside air

the level of oxygen inside the mine is lower

[1]

- 3 Hydrogen peroxide solution,  $\text{H}_2\text{O}_2$ , decomposes slowly in the absence of a catalyst. Oxygen and water are formed.

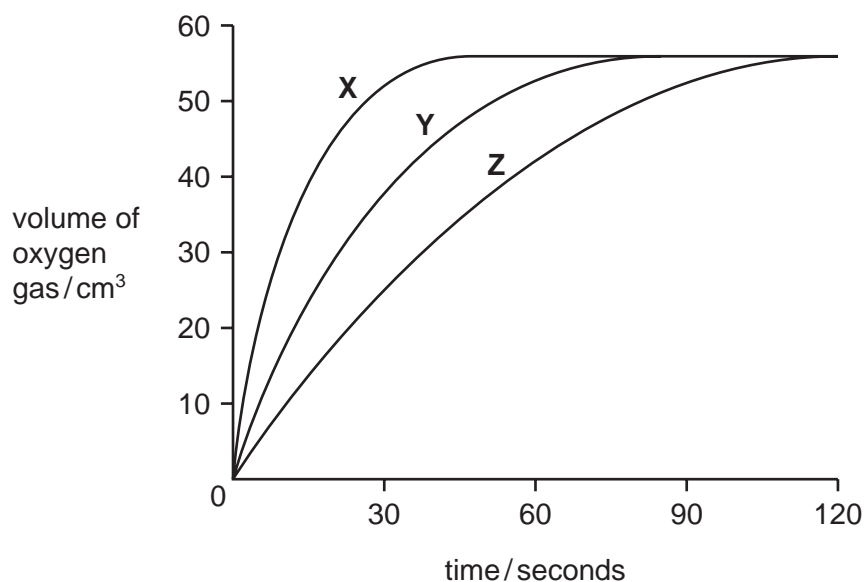


- (a) Draw a diagram of the apparatus you could use to investigate the speed of this reaction.

You must label your diagram.

[3]

- (b) Catalyst **X** was added to  $50\text{cm}^3$  of hydrogen peroxide solution at  $20^\circ\text{C}$  and the amount of oxygen given off was recorded over a two minute period. The experiment was repeated with the same amounts of catalyst **Y** and catalyst **Z**. Apart from the type of catalyst, all conditions were kept the same in the three experiments. A graph of the results is shown below.



- (i) What is a catalyst?

..... [1]

- (ii) Which catalyst, **X**, **Y** or **Z**, produced oxygen gas the fastest?  
Explain your answer.

.....  
..... [2]

- (iii) Why is the final amount of oxygen gas the same in each experiment?

.....  
..... [1]

- (iv) Many transition metals and their oxides are good catalysts.  
State **two** other properties of transition metals which are not shown by other metals.

.....  
..... [2]

- (c) The experiment with catalyst **Z** was repeated at 40°C.  
All other conditions were kept the same.  
The speed of the reaction increased.  
Explain why, using ideas about particles.

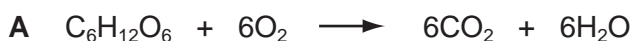
.....  
.....  
..... [2]

- (d) Some enzymes also catalyse the decomposition of hydrogen peroxide.

- (i) State one difference between an enzyme and an inorganic catalyst such as a transition metal.

..... [1]

- (ii) Enzymes are also responsible for fermentation reactions.  
Which one of the following equations **A**, **B**, **C** or **D** describes fermentation?



..... [1]



4 The list shows some oxides.

calcium oxide  
magnesium oxide  
nitrogen dioxide  
sodium oxide  
sulphur dioxide

(a) From this list choose **two** oxides which are basic.  
Give a reason for your answer.

.....  
..... [2]

(b) (i) Which **two** oxides from this list contribute to acid rain?

..... [2]

(ii) How do each of these oxides get into the atmosphere?

name of oxide .....  
source of oxide ..... [1]

name of oxide .....  
source of oxide ..... [1]

(c) Calcium oxide is manufactured from calcium carbonate.

(i) Complete the word equation for this reaction.

calcium carbonate  $\rightarrow$  calcium oxide + ..... [1]

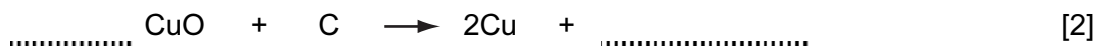
(ii) What condition is needed for this reaction to take place?

..... [1]

- (d) (i) Explain why calcium oxide and sodium oxide cannot be reduced by heating with carbon.

..... [1]

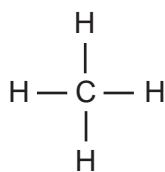
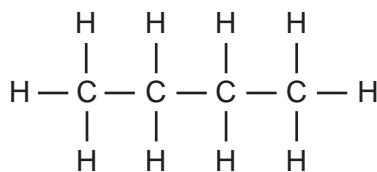
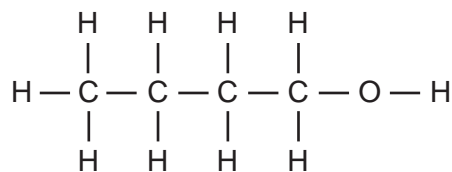
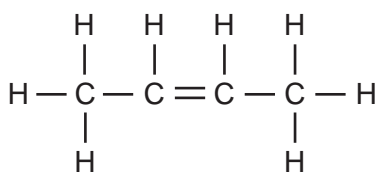
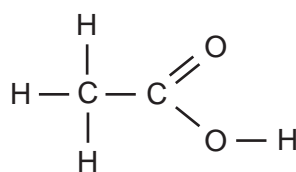
- (ii) Copper(II) oxide can be reduced by heating with carbon.  
Complete the equation for this reaction.



- (iii) What do you understand by the term *reduction*?

.....  
..... [1]

5 The structures of some organic compounds are shown below.

**A****B****C****D****E**

(a) Name compound **A**.

..... [1]

(b) Which **two** of the compounds **A** to **E** belong to the same homologous series?

..... [1]

(c) (i) Which one of the compounds **A** to **E** has the same functional group as ethanol?

..... [1]

(ii) Draw the structure of ethanol, showing all atoms and bonds.

[2]

(iii) Describe how ethanol is made in industry from ethene.

.....

..... [2]

(d) (i) Which one of the compounds **A** to **E** is an unsaturated hydrocarbon?

..... [1]

(ii) Describe a chemical test for an unsaturated hydrocarbon.

test .....

result ..... [2]

(e) Compound **E** is acidic.

(i) State the name of compound **E**.

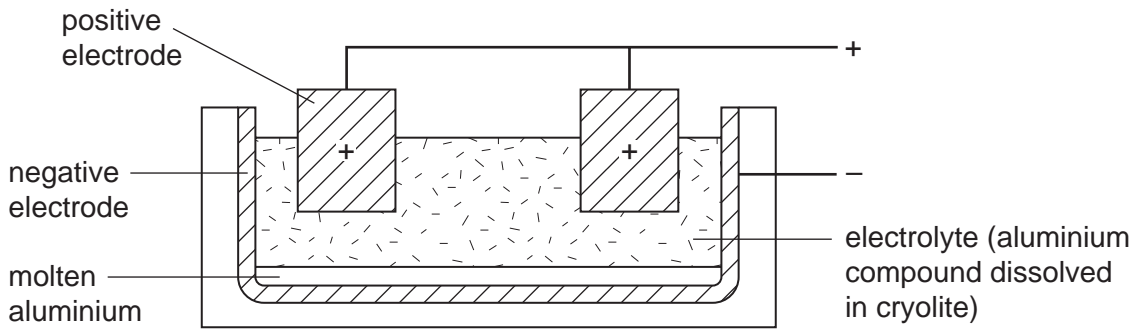
..... [1]

(ii) Describe a test to show that compound **E** is acidic.

test .....

result ..... [2]

6 The diagram shows an electrolysis cell used to extract aluminium.



(a) What compound of aluminium is used for the electrolyte?

..... [1]

(b) The electrolyte must be molten for the electrolysis to work. Explain why.

..... [1]

(c) (i) State the name of the substance used for the electrodes.

..... [1]

(ii) To which electrode do the aluminium ions move during electrolysis? Explain your answer.

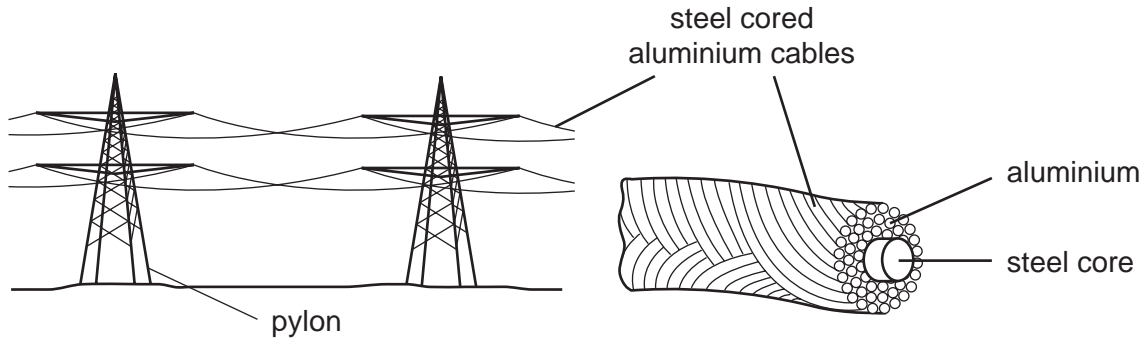
.....  
..... [2]

(d) Complete the following sentences about the molten electrolyte using words from the list below.

- |            |           |           |           |
|------------|-----------|-----------|-----------|
| bauxite    | chemical  | cryolite  | decreased |
| electrical | haematite | increased | light     |

The melting point of the electrolyte is ..... by adding  
..... This means that less ..... energy  
is needed to melt the electrolyte. [3]

(e) Aluminium is used in overhead power cables.



The table shows some properties of three metals which could be used for the power cables.

metal	relative electrical conductivity	density / grams per cm <sup>3</sup>	price / £ per kg	relative strength
aluminium	0.4	2.70	18	9
copper	0.7	8.92	15	30
steel	0.1	7.86	2.7	50

(i) Suggest why aluminium is used for overhead power cables rather than copper.

..... [1]

(ii) Suggest why steel is not used alone for overhead power cables.

..... [1]

(iii) Why is steel used as a core for overhead power cables?

..... [1]

(iv) Electrical insulators are used in parts of the pylons which carry the electrical cables. Which one of the following is an electrical insulator? Put a ring around the correct answer.

aluminium          ceramic          graphite          zinc          [1]

(f) Aluminium has many uses.

(i) Why is aluminium used for aircraft bodies?

..... [1]

(ii) Describe a test for aluminium ions.

test .....

result .....

..... [3]

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.

**DATA SHEET**  
**The Periodic Table of the Elements**

		Group										
I	II	III	IV	V	VI	VII	O					
		1 <b>H</b> Hydrogen 1										4 <b>He</b> Helium 2
7 <b>Li</b> Lithium 3	9 <b>Be</b> Beryllium 4											20 <b>Ne</b> Neon 10
23 <b>Na</b> Sodium 11	24 <b>Mg</b> Magnesium 12	11 <b>B</b> Boron 5	12 <b>C</b> Carbon 6	14 <b>N</b> Nitrogen 7	16 <b>O</b> Oxygen 8	19 <b>F</b> Fluorine 9					35.5 <b>Cl</b> Chlorine 17	
39 <b>K</b> Potassium 19	40 <b>Ca</b> Calcium 20	27 <b>Al</b> Aluminium 13	28 <b>Si</b> Silicon 14	31 <b>P</b> Phosphorus 15	32 <b>S</b> Sulphur 16	35.5 <b>Cl</b> Chlorine 17					40 <b>Ar</b> Argon 18	
85 <b>Rb</b> Rubidium 37	88 <b>Sr</b> Strontium 38	70 <b>Ga</b> Gallium 31	73 <b>Ge</b> Germanium 32	75 <b>As</b> Arsenic 33	79 <b>Se</b> Selenium 34	80 <b>Br</b> Bromine 35					84 <b>Kr</b> Krypton 36	
133 <b>Cs</b> Caesium 55	137 <b>Ba</b> Barium 56	115 <b>In</b> Indium 49	119 <b>Sn</b> Tin 50	122 <b>Sb</b> Antimony 51	128 <b>Te</b> Tellurium 52	127 <b>I</b> Iodine 53					131 <b>Xe</b> Xenon 54	
226 <b>Ra</b> Radium 88	227 <b>Ac</b> Actinium 89	204 <b>Tl</b> Thallium 81	207 <b>Pb</b> Lead 82	209 <b>Bi</b> Bismuth 83	201 <b>Hg</b> Mercury 80	84 <b>Po</b> Polonium 84					86 <b>Rn</b> Radon 86	

140 <b>Ce</b> Cerium 58	141 <b>Pr</b> Praseodymium 59	144 <b>Nd</b> Neodymium 60	150 <b>Sm</b> Samarium 62	152 <b>Eu</b> Europium 63	157 <b>Gd</b> Gadolinium 64	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	169 <b>Tm</b> Thulium 69	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71
232 <b>Th</b> Thorium 90	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	94 <b>Pu</b> Plutonium 94	95 <b>Am</b> Americium 95	96 <b>Cm</b> Curium 96	98 <b>Cf</b> Californium 98	99 <b>Es</b> Einsteinium 99	100 <b>Fm</b> Fermium 100	101 <b>Md</b> Mendelevium 101	102 <b>No</b> Nobelium 102	103 <b>Lr</b> Lawrencium 103

\*58-71 Lanthanoid series  
90-103 Actinoid series

a	X
b	b

Key  
a = relative atomic mass  
X = atomic symbol  
b = proton (atomic) number

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).