



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

--	--	--	--	--

CANDIDATE
NUMBER

--	--	--	--

COMBINED SCIENCE

0653/02

Paper 2 (Core)

May/June 2008

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.



READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
8	
9	
Total	

This document consists of **20** printed pages.



- 1 The Periodic Table shows all of the chemical elements arranged into groups and periods.

Fig. 1.1 shows part of the Periodic Table. The letters in this table are **not** the normal chemical symbols of the elements.

For
Examiner's
Use

	I	II									III	IV	V	VI	VII	0
1																A
2	F															E
3	C										H					
4	G															D

Fig. 1.1

- (a) Complete the statements below using letters, chosen from **A** to **H**, which refer to elements in Fig. 1.1. Letters may be used once, more than once or not at all.

- The element shown as letter is an alkali metal in period 3.
- The element shown as letter is the noble gas with the lowest density.
- The three elements shown as letters,, and have very similar chemical properties to each other.
- The element shown as letter is sometimes used as a catalyst. [4]

- (b) The elements sodium and sulphur are both oxidised when they burn in air to produce sodium oxide and sulphur dioxide respectively.

- (i) Explain the meaning of the term *oxidised*.

.....
..... [1]

(ii) Sodium oxide reacts with water to form solution **P**.

Sulphur dioxide reacts with water to form solution **Q**.

Predict and explain the colour of Universal Indicator solution when added to **P** and **Q**.

colour in **P**

explanation

.....

colour in **Q**

explanation

..... [4]

(iii) Name the type of chemical reaction which occurs when solution **P** is added to solution **Q**.

..... [1]

2 Fig. 2.1 shows the structure of the human thorax (seen from the front).

For
Examiner's
Use

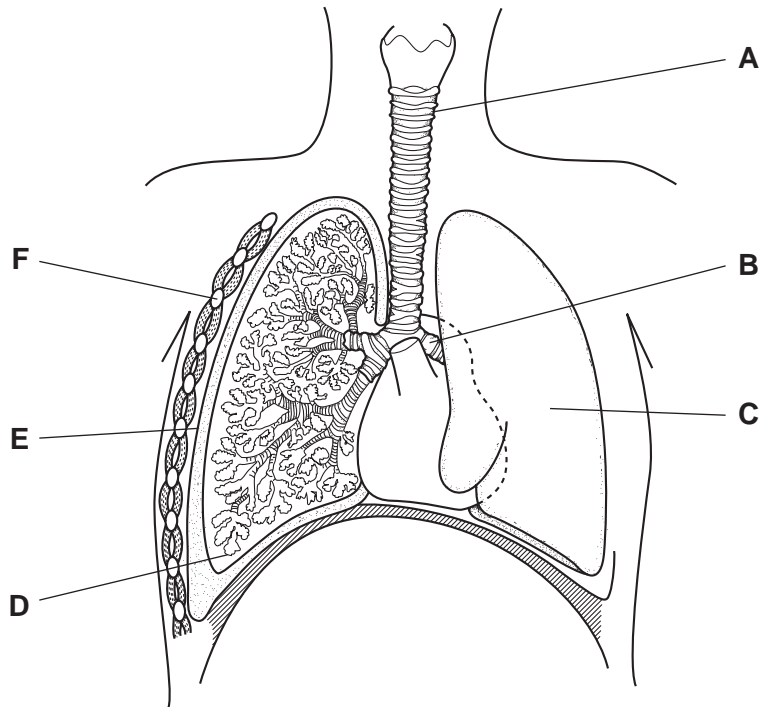


Fig. 2.1

(a) Give the **letter** of each of the following structures.

(i) the left bronchus

(ii) a pleural membrane

(iii) a place where there are goblet cells and cilia [3]

(b) Gas exchange takes place in the alveoli. When a person smokes for a number of years, the walls of the alveoli start to break down. This is called emphysema.

(i) Name the process by which molecules of oxygen pass into the blood from the alveoli.

..... [1]

(ii) Explain why emphysema makes it more difficult for oxygen to get into the blood.

.....

 [2]

(c) Oxygen is transported around the body in red blood cells. Fig. 2.2 is a diagram of a group of red blood cells.

For
Examiner's
Use

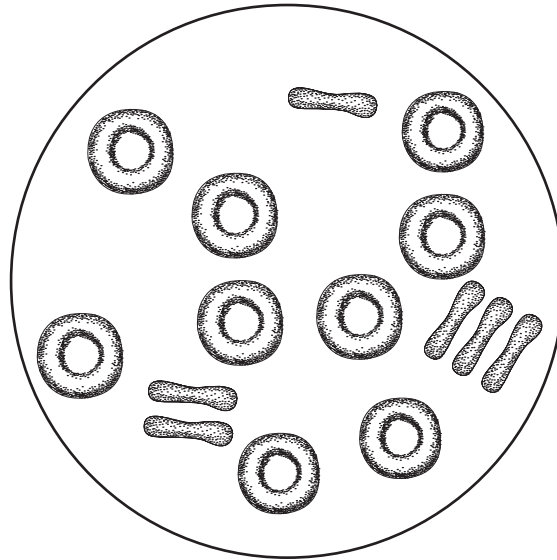


Fig. 2.2

(i) State **one** difference, apart from their colour, between the appearance of red blood cells and white blood cells.

..... [1]

(ii) What makes red blood cells look red?

..... [1]

(d) Explain why body cells need a constant supply of oxygen.

.....
.....
..... [2]

3 A man drives a golf ball with his club and it flies through the air for nearly 200 metres.

(a) (i) State the form of energy given to the ball by the club when the ball is hit.

..... [1]

(ii) State the type of energy gained by the ball as it rises into the air after being hit.

..... [1]

(b) As the golfer moves around the course in a golf cart, his movement is measured. The measurements are plotted on the graph in Fig. 3.1.

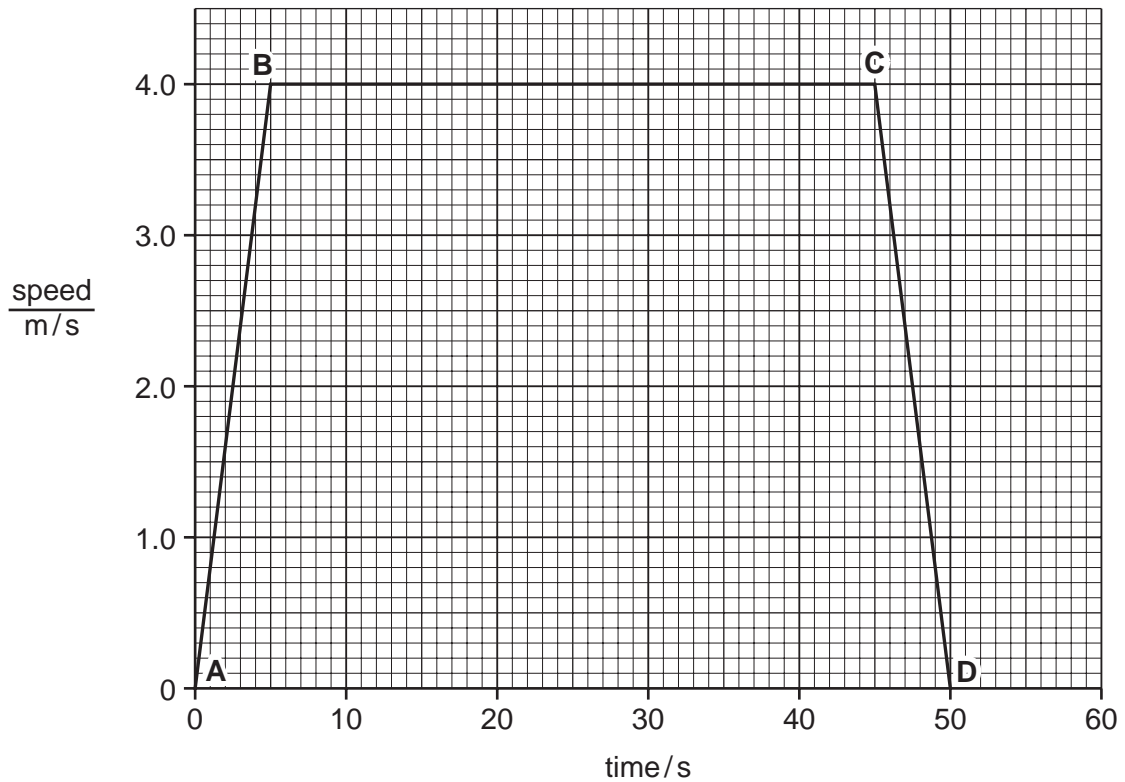


Fig. 3.1

(i) Describe what is happening between

A – B

.....

B – C

.....

C – D

..... [3]

(ii) What is the speed of the cart after 3 seconds?

..... m/s [1]

- (c) The golfer hits the ball along the ground. It travels 6 m in 3 s.

Calculate the average speed of the ball.

State the formula that you use and show your working.

formula

working

..... m/s [2]

- (d) The golfer's bag of clubs has a mass of 6 kg.

- (i) Calculate the weight of the bag of clubs.

Assume that the gravitational field strength on Earth is 10 N/kg.

..... N [1]

- (ii) Calculate the work done by the golfer when the bag is lifted 0.5 m.

State the formula that you use and show your working.

formula

working

..... J [2]

4 Kerosene is a mixture of hydrocarbons used as a fuel for aircraft and for lighting and cooking.

For
Examiner's
Use

(a) Kerosene is obtained from petroleum (crude oil) and is a liquid which boils in the range 150 °C – 200 °C.

(i) Name the process used to separate kerosene from petroleum.

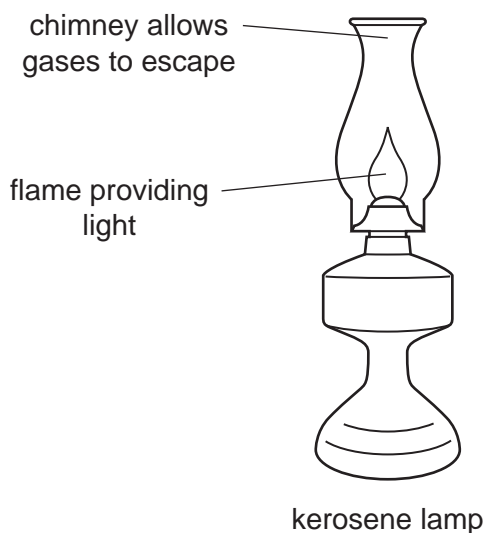
..... [1]

(ii) State the important difference between the various compounds in petroleum which enables them to be separated by the process you have named in (i).

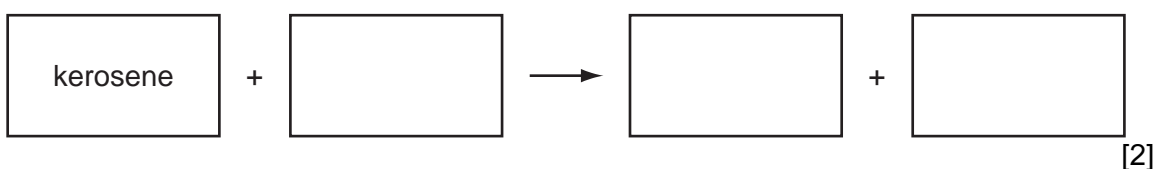
..... [1]

(b) The light from a kerosene lamp is provided by the flame produced when kerosene burns in air.

The lamp must be carefully designed and operated to ensure that most of the kerosene undergoes complete combustion.



(i) Complete the **word** chemical equation for the complete combustion of kerosene.



- (ii) Describe **one** observation which shows that the reaction occurring in the kerosene lamp is exothermic.

.....
.....
..... [1]

- (c) The full chemical symbol for carbon is shown below.



Draw a diagram of a carbon atom. Label the nucleus and show the full electron configuration.

[2]

- 5 Fig. 5.1 shows the quantity of carbon dioxide that was emitted to the atmosphere by a large industrial company, between 2000 and 2005.

For
Examiner's
Use

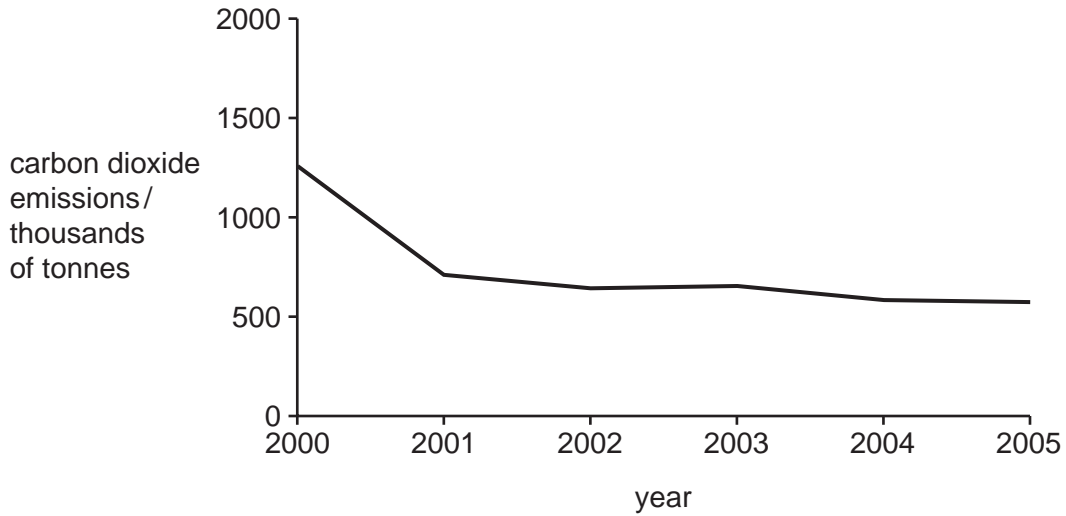


Fig. 5.1

- (a) Describe how the company's carbon dioxide emissions changed between 2000 and 2005.

.....
.....
..... [2]

- (b) The company stated that these carbon dioxide emissions included those relating to the electricity that it used.

Explain how using electricity can be responsible for emissions of carbon dioxide.

.....
.....
..... [2]

- (c) Apart from using less electricity, suggest **one** other way that the company could reduce its carbon dioxide emissions.

.....
..... [1]

- (d) In 1997, at a meeting in the city of Kyoto in Japan, many countries in the world signed an agreement to reduce their emissions of carbon dioxide. The agreement came into force in 2005.

For
Examiner's
Use

Explain why we need to reduce emissions of carbon dioxide.

.....
.....
..... [2]

- (e) Tropical rainforests can help to combat rising levels of carbon dioxide, because they take it from the air and use it in photosynthesis.

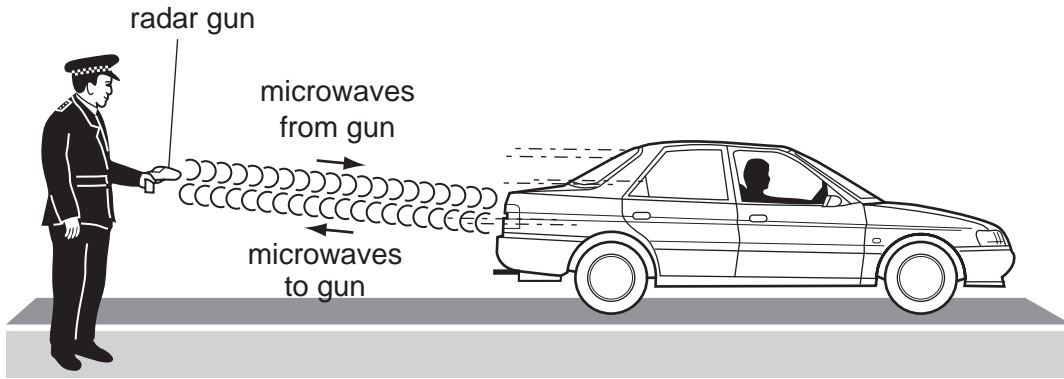
Describe **one** other reason why we should try to conserve tropical rainforests.

.....
.....
..... [2]

- 6 (a) A policeman is using a radar gun to measure the speed of a car.

The radar gun emits microwaves which hit the moving car and bounce back to a receiver in the radar gun.

A computer in the radar gun calculates the speed of the car.



- (i) What type of waves are microwaves?

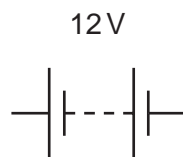
..... [1]

- (ii) The waves bounce off the car back towards the radar gun. What is this process called?

..... [1]

- (b) A car has two headlamps and two rear lamps. All four lamps are connected in parallel with each other across a 12V battery.

- (i) Complete the circuit diagram below to show how the four lamps are connected to the battery. Include one switch in your circuit which will control all four lamps.



[2]

(ii) If the filament in one lamp breaks, the other three stay lit. Explain why this happens.

.....
..... [1]

For
Examiner's
Use

(c) Fig. 6.1 shows a spring. The spring is 10 cm long. A metal nut is hung on the spring and the length is now 13 cm.

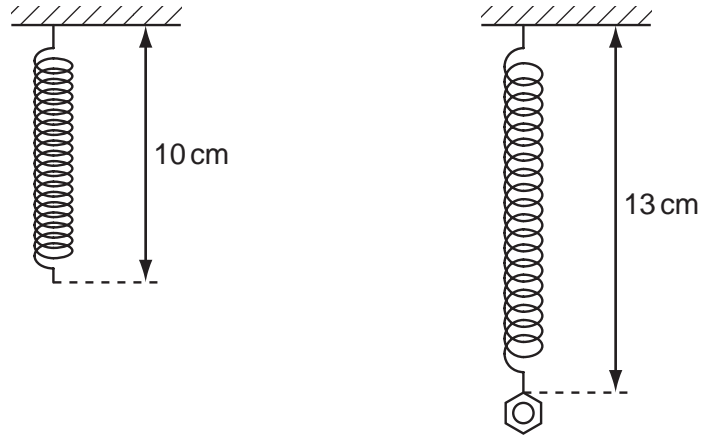


Fig. 6.1

Calculate the length of the spring if 3 **more** identical nuts are hung on the spring.
Show your working.

.....cm [2]

7 (a) Draw lines to link each term to its definition.

term	definition
cell membrane	a green pigment found in some plant cells, which absorbs energy from sunlight
chlorophyll	a partially permeable layer surrounding a cell
cell wall	a fully permeable layer surrounding a plant cell
chloroplast	an organelle found in some plant cells, where photosynthesis takes place

[3]

(b) Plant leaves often contain starch, which is produced during photosynthesis.

Describe how the starch is produced.

.....

.....

.....

..... [2]

- (c) Fig. 7.1 shows one of the ways in which a plant called *Bryophyllum* reproduces. It grows new plantlets from its leaves.

For
Examiner's
Use



Fig. 7.1

- (i) Name the type of reproduction that is taking place.

..... [1]

- (ii) The new plants that are produced are clones of the parent plant.

Explain what is meant by the term *clone*.

.....

 [2]

8 (a) A student wrote down some properties of alpha, beta and gamma radiations.

Draw a line from each property to the correct radiation.

stopped by paper	alpha
contains negatively charged particles	
passes through several centimetres of lead	beta
passes through paper but stopped by a few millimetres of aluminium	
has no mass	gamma

[3]

(b) (i) Gamma radiation can be used to sterilise surgical instruments. What property of gamma radiation makes it suitable for this purpose?

..... [1]

(ii) State **one** other use for radiation from a radioactive source.

.....
..... [1]

(c) In an experiment a radiation detector was set up and used to measure background radiation. The background radiation in the laboratory was found to be 40 counts per minute.

(i) What is *background radiation*?

.....
..... [1]

(ii) State **one** source of background radiation.

..... [1]

(iii) A radioactive source was placed near the detector and a reading of 1200 counts per minute was recorded. What was the count rate of the radioactive source?

..... counts per minute [1]

- 9 Fig. 9.1 shows apparatus which can be used to reduce copper oxide to copper.

Copper oxide is a black powder and during the reaction metallic copper forms inside the reaction tube.

For
Examiner's
Use

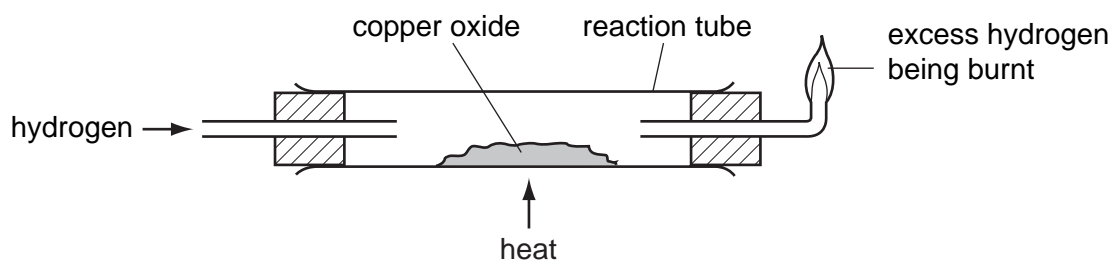
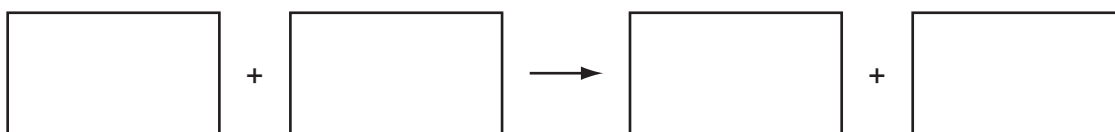


Fig. 9.1

- (a) (i) Select from the list of substances below to complete the word equation for the reaction in Fig. 9.1.

air	copper	copper oxide
hydrogen	oxygen	water



[1]

- (ii) Describe **one** piece of evidence which would show that copper had been formed in this reaction.

.....

..... [1]

- (b) When a student carried out the reaction in Fig. 9.1 she realised the material left inside the reaction tube was a mixture of metallic copper and unreacted copper oxide.

In order to separate the metallic copper, she stirred the material from the reaction tube with warm dilute sulphuric acid for several minutes. She then filtered the mixture as shown in Fig. 9.2.

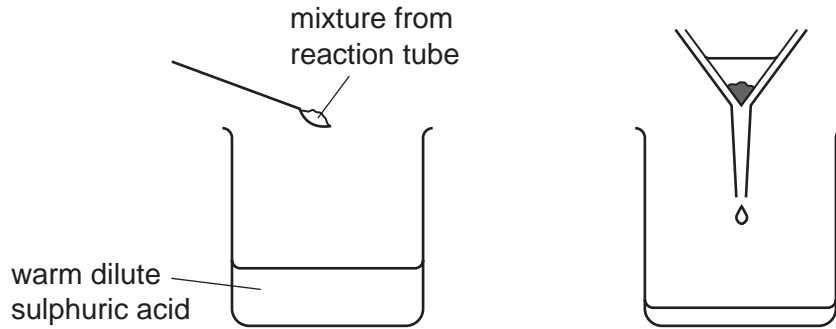


Fig. 9.2

- (i) Name the copper compound formed when sulphuric acid reacts with copper oxide.
 [1]

- (ii) The copper compound you have named in (i) is soluble.

Explain why the method shown in Fig. 9.2 is successful in separating metallic copper from the original mixture of copper and copper oxide.

.....

 [2]

- (c) Copper oxide is a compound of a metal and a non-metal.

- (i) Name the type of chemical bonding in copper oxide.
 [1]

- (ii) Explain why there is a strong force of attraction between the copper and oxide particles in copper oxide.

.....

 [2]

(d) Metallic copper can also be obtained by electrolysis.

For
Examiner's
Use

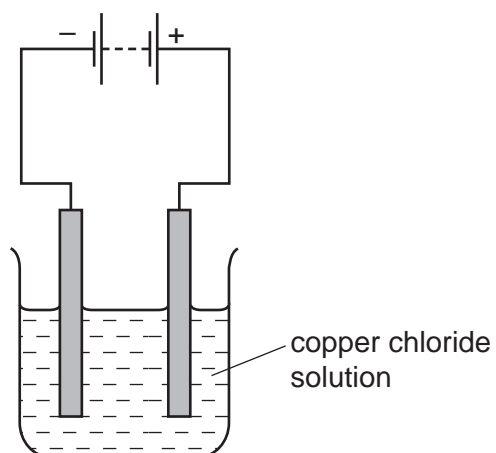


Fig. 9.3

Describe what would be seen at each of the electrodes when the electrolysis shown in Fig. 9.3 is carried out.

at the positive electrode

.....

at the negative electrode

.....

[2]

DATA SHEET
The Periodic Table of the Elements

		Group															
		I	II	III	IV	V	VI	VII	VIII	IX	X						
		1 H Hydrogen 1															
7	9	Li Lithium 3	Be Beryllium 4														
23	24	Na Sodium 11	Mg Magnesium 12														
39	40	K Potassium 19	Ca Calcium 20	51	52	55	56	59	59	64	65	70	73	75	79	80	84
85	88	Rb Rubidium 37	Sr Strontium 38	V Vanadium 23	Ti Titanium 22	Mn Manganese 25	Fe Iron 26	Co Cobalt 27	Ni Nickel 28	Cu Copper 29	Zn Zinc 30	Ga Gallium 31	Ge Germanium 32	As Arsenic 33	Se Selenium 34	Br Bromine 35	Kr Krypton 36
133	137	Cs Caesium 55	Ba Barium 56	93	91	96	101	103	106	108	112	115	119	122	128	127	131
		Fr Francium 87	Ra Radium 88	Nb Niobium 41	Zr Zirconium 40	Tc Technetium 43	Ru Ruthenium 44	Rh Rhodium 45	Pd Palladium 46	Ag Silver 47	Cd Cadmium 48	In Indium 49	Sn Tin 50	Sb Antimony 51	Te Tellurium 52	I Iodine 53	Xe Xenon 54
				181	178	186	190	192	195	197	201	204	207	209	210	210	210
				Ta Tantalum 73	Hf Hafnium 72	Re Rhenium 75	Os Osmium 76	Ir Iridium 77	Pt Platinum 78	Au Gold 79	Hg Mercury 80	Tl Thallium 81	Pb Lead 82	Bi Bismuth 83	Po Polonium 84	At Astatine 85	Rn Radon 86
				140	141	144	150	152	157	159	162	165	167	169	173	175	
				Ce Cerium 58	Pr Praseodymium 59	Nd Neodymium 60	Pm Promethium 61	Sm Samarium 62	Eu Europium 63	Gd Gadolinium 64	Tb Terbium 65	Dy Dysprosium 66	Ho Holmium 67	Er Erbium 68	Tm Thulium 69	Yb Ytterbium 70	Lu Lutetium 71
				232	238	238	238	238	238	238	238	238	238	238	238	238	238
				Th Thorium 90	Pa Protactinium 91	U Uranium 92	Np Neptunium 93	Pu Plutonium 94	Am Americium 95	Cm Curium 96	Bk Berkelium 97	Cf Californium 98	Es Einsteinium 99	Fm Fermium 100	Md Mendelevium 101	No Nobelium 102	Lr Lawrencium 103

*58-71 Lanthanoid series
†90-103 Actinoid series

a = relative atomic mass

X = atomic symbol

b = proton (atomic) number

Key

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.