

730

## UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

| CANDIDATE<br>NAME |                       |
|-------------------|-----------------------|
| CENTRE<br>NUMBER  | CANDIDATE<br>NUMBER   |
| COMBINED SC       | 0653/23               |
| Paper 2 (Core)    | October/November 2012 |
|                   | 1 hour 15 minutes     |

Candidates answer on the Question Paper.

No Additional Materials are required.

## READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in. Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions. A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together. The number of marks is given in brackets [ ] at the end of each question or part question.

| For Examiner's Use |  |  |
|--------------------|--|--|
| 1                  |  |  |
| 2                  |  |  |
| 3                  |  |  |
| 4                  |  |  |
| 5                  |  |  |
| 6                  |  |  |
| 7                  |  |  |
| 8                  |  |  |
| 9                  |  |  |
| Total              |  |  |

This document consists of 20 printed pages.



1 Flowers are organs in which sexual reproduction takes place. For Examiner's Use (a) (i) Complete the definition of sexual reproduction. Use words from the list. dissimilar female haploid identical ovary sperm zygote Sexual reproduction is the process involving the fusion of nuclei to form a diploid and the production of genetically offspring. [3] (ii) State the scientific term for the fusion of two nuclei. [1] ..... (b) Fig. 1.1 shows a section through a flower. В D. Fig. 1.1 (i) Name the parts labelled A and B. Α [2] В \_\_\_\_\_ (ii) State the letter of the part in which the male gametes are produced, a zygote is produced. [2] .....

(c) After pollination, seeds are produced. A student set up an experiment to investigate the conditions needed for the germination of lettuce seeds.

He placed five lettuce seeds on cotton wool in each of five test-tubes. Fig. 1.2 shows the conditions present in each tube.

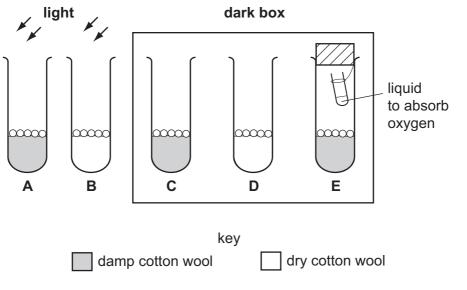




Table 1.1 shows his results.

| Table | 1 | .1 |
|-------|---|----|
|-------|---|----|

| tube | conditions |        |       | number of seeds that germinated |
|------|------------|--------|-------|---------------------------------|
| Α    | water      | oxygen | light | 5                               |
| В    | no water   | oxygen | light | 0                               |
| С    |            |        |       | 5                               |
| D    |            |        |       | 0                               |
| E    |            |        |       | 0                               |

- (i) Complete Table 1.1 to show the conditions present in each tube. Tubes A and B have been done for you. [2]
- (ii) What conclusions can the student make from these results?

[3]

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(a) (i) State the percentages of nitrogen and oxygen in the air. 2 Examiner's Use nitrogen [2] oxygen ..... (ii) During a thunderstorm, energy from lightning causes nitrogen and oxygen to combine to form nitric oxide. Explain why nitrogen is an example of an *element* and nitric oxide is an example of a compound. [2] (iii) Nitric oxide has the chemical formula, NO. Explain what is meant by this formula. ..... ......[2] (iv) What name is given to the type of chemical reaction that occurs when oxygen bonds to another element? .....[1] (b) When magnesium burns in air, a white solid is formed. This white solid contains magnesium oxide, MgO. (i) Name the type of chemical bonding in magnesium oxide. Explain your answer. type of chemical bonding explanation ..... [2]

For

(ii) A student burned some magnesium in air and then added the white solid formed to water. Examiner's

She tested the solution with Universal (full range) Indicator and found that the pH was 9.

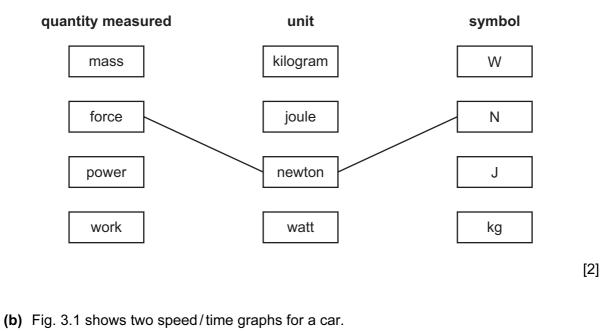
State a conclusion that the student can draw from this observation.

..... ......[1] For

Use

**3** (a) Draw lines to connect each quantity measured to its correct unit and symbol.

One has been done for you.



speed graph A graph B graph B



Describe the motion of the car in graph **A**, \_[2]

For Examiner's Use

| (c) | The  | e car travels at 20 m/s for 90 seconds.  |     | For               |
|-----|------|--|-----|-------------------|
|     | Cal  | culate the distance covered.   |     | Examiner's<br>Use |
|     | Sta  | te the formula that you use and show your working.                               |     |                   |
|     |      | formula used   |     |                   |
|     |      | working  |     |                   |
|     |      |  |     |                   |
|     |      |  |     |                   |
|     |      |  |     |                   |
|     |      | m  | [2] |                   |
| (d) | On   | e of the car's headlamps has a current of 2A, when the voltage across it is 12V. |     |                   |
|     | (i)  | Show that the resistance of the headlamp is $6 \Omega$ .                         |     |                   |
|     |      | State the formula that you use and show your working.                            |     |                   |
|     |      | formula used   |     |                   |
|     |      | working  |     |                   |
|     |      |  |     |                   |
|     |      |  |     |                   |
|     |      |  | [2] |                   |
|     | (ii) | The car has two of these identical headlamps connected in series.                |     |                   |
|     |      | Calculate the total resistance of these two headlamps.                           |     |                   |
|     |      | State the formula that you use and show your working.                            |     |                   |
|     |      | formula used   |     |                   |
|     |      | working  |     |                   |
|     |      |  |     |                   |
|     |      |  |     |                   |
|     |      |  |     |                   |
|     |      | Ω  | [2] |                   |

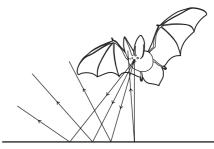
| 4 | Many bats are predators that fly at night. They eat moths and other insects.  |                   |             |                    |                 | For<br>Examiner's<br>Use |     |  |
|---|---|-------------------|-------------|--------------------|-----------------|--------------------------|-----|--|
|   | (a) Underline the <b>two</b> words that describe the position of a bat in a food chain.   |                   |             |                    |                 |                          |     |  |
|   | C   | arnivore          | consumer    | herbivore          | producer        |                          | [1] |  |
|   |   | s emit ultrasound |             | upper too bigh for | a human ta haar |                          |     |  |
|   | <ul> <li>(i) Ultrasound is sound that has a frequency too high for a human to hear.</li> <li>Suggest a frequency for the ultrasound emitted by bats.</li> </ul> |                   |             | Hz                 | [1]             |                          |     |  |
|   | (ii) Underline the <b>one</b> word that correctly describes an ultrasound wave.   |                   |             |                    |                 |                          |     |  |
|   |   | electroma         | agnetic lon | gitudinal          | transverse      |                          | [1] |  |

(c) Bats use echo location to detect objects around them.

The reflected ultrasound waves are detected by special cells in the bat's head.

Fig. 4.1 shows how ultrasound waves are reflected from a rough surface and from a smooth surface. The arrows show the direction in which the sound waves travel.

rough surface



smooth surface

Fig. 4.1

(i) Use the information in Fig. 4.1 to describe what happens to the ultrasound waves when they hit

| a rough surface,  | <br> |
|-------------------|------|
|                   |      |
| a smooth surface. |      |
|                   | [1]  |
|                   |      |

(ii) Suggest how the bat can tell if it is flying over a rough surface or a smooth surface, even when it is completely dark.
[1]
(d) Many kinds of bat live in trees in forests.
List three ways in which deforestation can damage the environment.
1
2
3
[3]

For

Examiner's Use 5 (a) In many countries, river water is collected and treated to make it safe for humans to For drink. Examiner's Use State and explain which two of the processes shown below are used to treat river water so that it becomes safe to drink. chlorination crystallisation filtration evaporation first process ..... reason why this process is carried out ..... second process reason why this process is carried out [4]

(b) Fig. 5.1 shows chromatography being used by a student to investigate mixtures of dyes (coloured compounds) used to colour sweets.

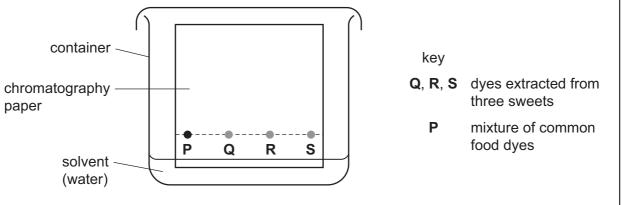
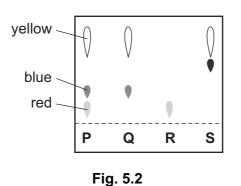
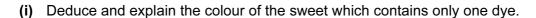




Fig. 5.2 shows the appearance of the chromatography paper after several minutes.





|      | colour                  |  |
|------|-------------------------|--|
|      | explanation             |  |
|      |                         | [2]  |
| (ii) | State which the mixture | sweet contained a dye which was <b>not</b> one of the common food dyes in <b>P</b> . |

[1]

For

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Use

6 (a) Fig. 6.1 shows a washing machine.





Complete the sentence below using  $\ensuremath{\textit{two}}$  of the words in the list.

|         | heat            | kinetic           | light          | potential            | sound   |     |
|---------|-----------------|-------------------|----------------|----------------------|---------|-----|
| Av      | vashing machine | e is designed to  | transform ele  | ectrical energy into |         |     |
|         |                 | ener              | gy and         |                      | energy. | [2] |
| (b) (i) | Some of the wa  | ater inside the v | vashing mach   | nine evaporates.     |         |     |
|         | Explain the pro | cess of evapora   | ation in terms | of particles.        |         |     |
|         |                 |                   |                |                      |         |     |
|         |                 |                   |                |                      |         |     |
|         |                 |                   |                |                      |         |     |
|         |                 |                   |                |                      |         | [2] |
| (ii)    | Explain why ev  | aporation has a   | a cooling effe | ct.                  |         |     |
|         |                 |                   |                |                      |         |     |
|         |                 |                   |                |                      |         | [1] |

(c) The casing of the washing machine is a solid. The water used in it is a liquid.

Complete the diagrams below to show the arrangement of particles in a solid and in a liquid.

solid liquid

[2]

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(d) Before buying a washing machine, a person may research several types to find out which washing machine has the greatest energy efficiency.

Explain the meaning of the term efficiency.

[1]

7 (a) Fig. 7.1 shows two human teeth.

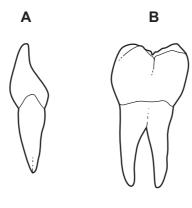


Fig. 7.1

(i) Name the **two** types of teeth shown in Fig. 7.1.

tooth A \_\_\_\_\_\_

(ii) Explain how tooth **B** helps to digest a food such as bread.

[2]

(b) For each part of the digestive system in the list below, tick (✓) the correct function or functions.

| part            | ingestion | digestion | absorption |
|-----------------|-----------|-----------|------------|
| mouth           |           |           |            |
| stomach         |           |           |            |
| small intestine |           |           |            |

[3]

[2]

For Examiner's Use Please turn over for Question 8.

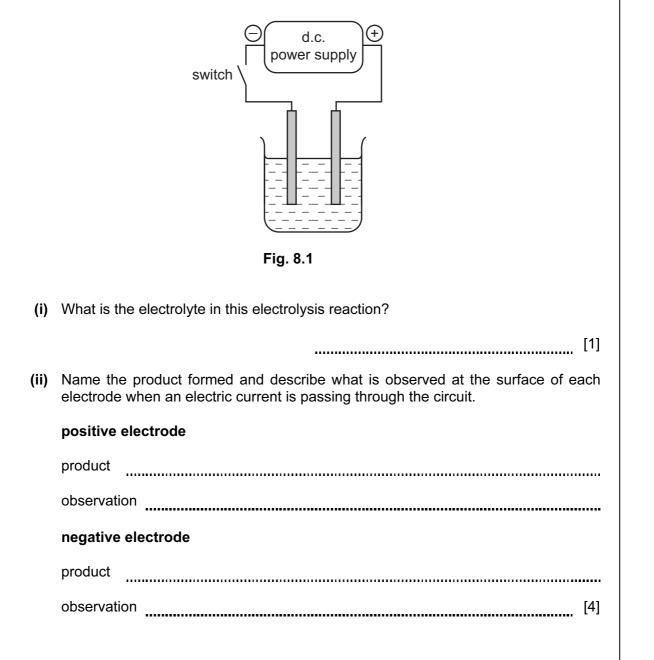
15

Metallic copper is a very important material that has been extracted from copper 8 For compounds for thousands of years. Examiner's Use (a) (i) The wires used in many electrical devices are made from copper. State the two properties of metals such as copper, that make them suitable for making electrical wires. ..... 1 [2] 2 (ii) Copper wires are connected to the mains electrical supply using brass plugs. Brass is an alloy. copper wire brass plug -Explain the meaning of the term alloy and state one difference in the physical properties of brass compared to copper. meaning of alloy ..... difference in physical property [2] (iii) One of the processes used in the extraction of copper involves heating copper(I) sulfide in air. One of the reactions that occurs is between copper(I) sulfide and oxygen. This reaction also produces sulfur dioxide. Construct the word chemical equation for this reaction. ......[1]

(b) Copper may also be formed by the electrolysis of an aqueous solution of copper chloride using electrodes made of graphite (carbon).

For Examiner's Use

Fig. 8.1 shows a laboratory apparatus a student used to carry out this electrolysis reaction.



**9** (a) Complete Table 9.1 to show the circuit symbol for each of the named components.

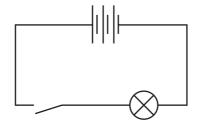
| component         | symbol |
|-------------------|--------|
| ammeter           |        |
| fuse              |        |
| variable resistor |        |

## Table 9.1

For Examiner's Use

[3]

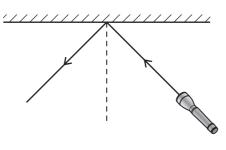
(b) Fig. 9.1 shows an electrical circuit for a torch (flashlight).





| (i)  | How many cells are fitted in the torch?                                   | [1] |
|------|---|-----|
| (ii) | A voltmeter is used to check the voltage across the light bulb.           |     |
|      | Draw the symbol for the voltmeter in the correct position on the circuit. | [1] |

(c) A single ray of light from a torch is shone onto a mirror as shown in Fig. 9.2.





- (i) On Fig. 9.2, label the angle of incidence and angle of reflection.
- (ii) The angle of incidence =  $45^{\circ}$ .

Write down the value of the angle of reflection.

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[1]

[1]

.....

|       | 0 | <sup>4</sup> <b>He</b> | Helium<br>2   | 20    | Ne     | Neon<br>10     | 40            | Ar | Argon<br>18     | 84               | Kr | Krypton<br>36   | 131           | Xe  | Xenon<br>54      |                | Rn  | Radon<br>86       |               |                      | 175                      | Lu                      | Lutetium<br>71     |                          | ۲                 | Lawrencium<br>103          |                |
|-------|---|------------------------|---------------|-------|--------|----------------|---------------|----|-----------------|------------------|----|-----------------|---------------|-----|------------------|----------------|-----|-------------------|---------------|----------------------|--------------------------|-------------------------|--------------------|--------------------------|-------------------|----------------------------|----------------|
| Group |   |                        |               | 19    | Ľ      | Fluorine<br>9  | 35.5          | Cl | Chlorine<br>17  | 80               | Ŗ  | Bromine<br>35   | 127           | н   | lodine<br>53     |                | At  | Astatine<br>85    |               |                      | 173                      | Υb                      | Ytterbium<br>70    |                          | No                | Nobelium                   |                |
|       | > |                        |               | 16    | 0      | Oxygen<br>8    | 32            | S  | Sulfur<br>16    | 79               | Se | Selenium<br>34  | 128           | Te  | Tellurium<br>52  |                | Ро  | Polonium<br>84    |               |                      | 169                      | Tm                      | Thulium<br>69      |                          | Md                | Mendelevium                |                |
|       | > |                        | -             |       | 14     | z              | Nitrogen<br>7 | 31 | ٩               | Phosphorus<br>15 | 75 | As              | Arsenic<br>33 | 122 | Sb               | Antimony<br>51 | 209 | Bi                | Bismuth<br>83 |                      |                          | 167                     | ц                  | Erbium<br>68             |                   | Еm                         | Fermium<br>100 |
|       | ≥ |                        |               | 12    | с<br>С | Carbon<br>6    | 28            | Si | Silicon<br>14   | 73               | Ge | Germanium<br>32 | 119           | Sn  | Tin<br>50        | 207            | Pb  | Lead<br>82        |               |                      | 165                      | Ч                       | Holmium<br>67      |                          | Es                | Einsteinium<br>aa          |                |
|       | ≡ |                        |               | 5     | ß      | Boron<br>5     | 27            | ١V | Aluminium<br>13 | 70               | Ga | Gallium<br>31   | 115           | In  | Indium<br>49     | 204            | 11  | Thallium<br>81    |               |                      | 162                      | Dy                      | Dysprosium<br>66   |                          | ç                 | Californium<br>QR          |                |
|       |   |                        |               |       |        |                |               |    |                 |                  | Zn | Zinc<br>30      | 112           | Cd  | Cadmium<br>48    | 201            | Hg  | Mercury<br>80     |               |                      | 159                      | Tb                      | Terbium<br>65      |                          |                   | Berkelium<br>97            |                |
|       |   |                        |               |       |        |                |               |    |                 | 64               | Cu | Copper<br>29    | 108           | Ag  | Silver<br>47     | 197            | Au  | Gold<br>79        |               |                      | 157                      | Gd                      | Gadolinium<br>64   |                          | Cm                | Curium<br>GG               |                |
|       |   |                        |               |       |        |                |               |    |                 | 59               | ïZ | Nickel<br>28    | 106           | Pd  | Palladium<br>46  | 195            | F   | Platinum<br>78    |               |                      | 152                      | Eu                      | Europium<br>63     |                          | Am                | Americium                  |                |
|       |   |                        |               |       |        |                |               |    |                 | 59               | ပိ | Cobalt<br>27    | 103           | Rh  | Rhodium<br>45    | 192            | ŗ   | Iridium<br>77     |               |                      | 150                      | Sm                      | Samarium<br>62     |                          |                   | Plutonium<br>0.4           |                |
|       |   | - I                    | Hydrogen<br>1 |       |        |                |               |    |                 | 56               | Fe | lron<br>26      | 101           | Ru  | Ruthenium<br>44  | 190            | 0s  | Osmium<br>76      |               |                      |                          |                         | Promethium<br>61   |                          | Np                | Neptunium                  |                |
|       |   |                        |               |       |        |                |               |    |                 | 55               | Mn | Manganese<br>25 |               | Ц   | Technetium<br>43 | 186            | Re  | Rhenium<br>75     |               |                      | 144                      |                         | Neodymium<br>60    | 238                      |                   | Uranium                    |                |
|       |   |                        |               |       |        |                |               |    |                 | 52               | ບັ | Chromium<br>24  | 96            | Мо  | Molybdenum<br>42 | 184            | ≥   | Tungsten<br>74    |               |                      | 141                      | Pr                      | Praseodymium<br>59 |                          | Ра                | Protactinium<br>01         |                |
|       |   |                        |               |       |        |                |               |    |                 | 51               | >  | Vanadium<br>23  | 93            | qN  | Niobium<br>41    | 181            | Та  | Tantalum<br>73    |               |                      | 140                      | Ce                      | Cerium<br>58       | 232                      | Ч                 | Thorium                    |                |
|       |   |                        |               |       |        |                |               |    |                 | 48               | F  | Titanium<br>22  | 91            | Zr  | Zirconium<br>40  | 178            | Ηf  | Hafnium<br>72     |               |                      | _                        |                         |                    | nic mass                 | pol               | nic) number                |                |
|       |   |                        |               |       |        |                |               |    |                 | 45               | Sc | Scandium<br>21  | 89            | ≻   | Yttrium<br>39    | 139            | La  | Lanthanum<br>57 * | 227           | Ac<br>Actinium<br>89 | *58-71 Lanthanoid corioc | eries                   | 2222               | a = relative atomic mass | X = atomic symbol | b = proton (atomic) number |                |
|       |   | 1                      |               |       |        | m              | 24            | Mg | nesium          | 40               | Ca | Calcium         | 88            | Sr  | Strontium<br>3   | 137            |     | Barium            | 226           | Radium<br>Radium     | pioned                   | 190-113 Actinoid series | 20                 | <br>9                    | ×                 | -<br>P                     |                |
|       | = |                        |               | б<br> | Be     | Beryllium<br>4 | Ď,            | Σ  | Magnesium<br>12 |                  |    | 20 C            |               |     | 38               |                |     | 56                |               | 88                   |                          | Acti                    | 2                  | ŋ                        | ×                 |                            |                |

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